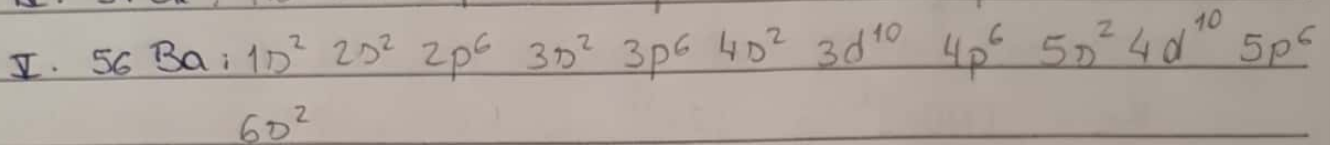
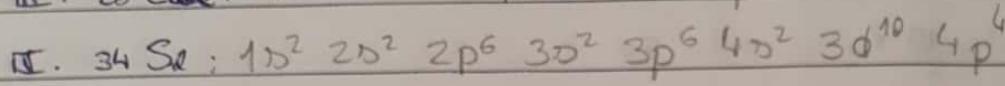
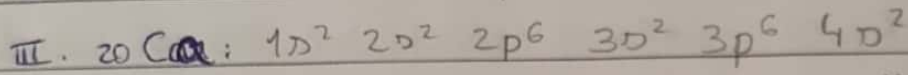
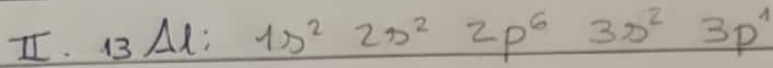
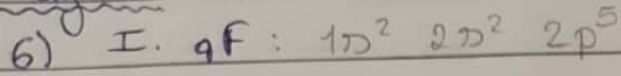


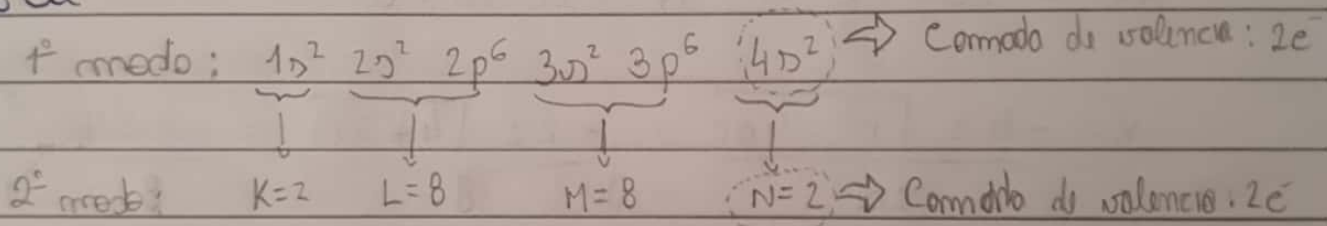
Resolução dos exercícios de EF9

Livro - Parte I.

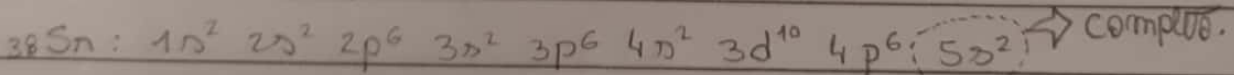
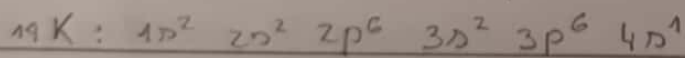
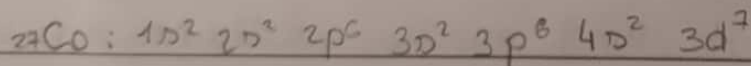
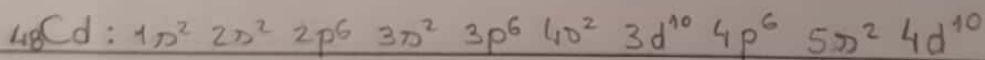
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8) $20Ca$



9) forma tradicional:

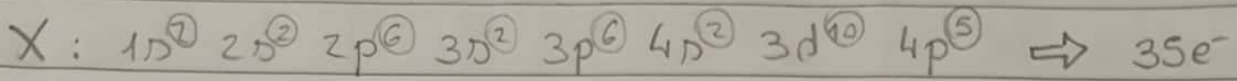


Subnível mais energético
é a camada de valência.

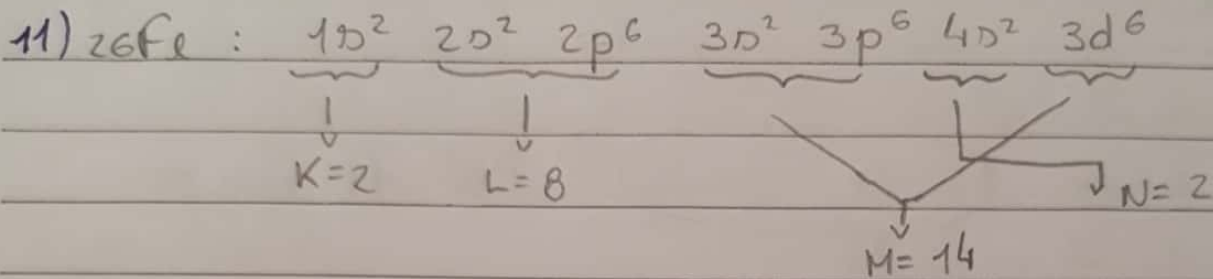
Nota 8

Os elementos que possuem o subnível "d" completo e tem ele como sendo o mais energético não se pertencem ao grupo 2 da Tabela, além do átomo de Hélio.

10) Substituímos mais energético \Rightarrow última da distribuição!



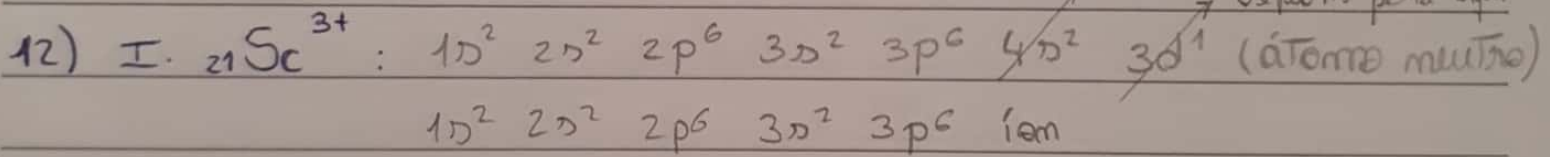
Supondo um átomo neutro, $35X$. Ou seja, $Z=35$.



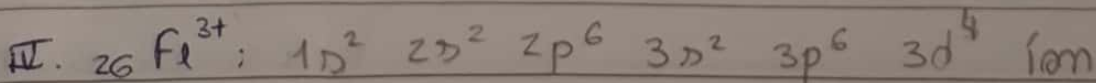
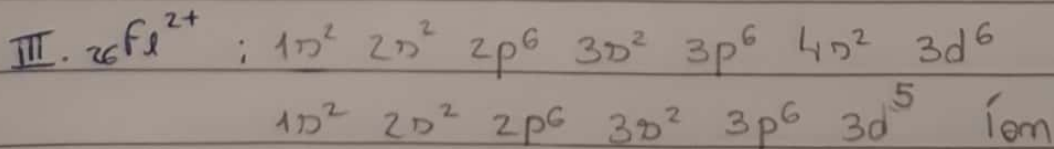
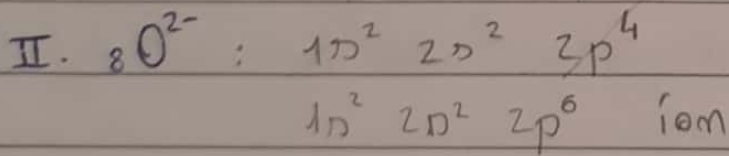
Então, 2-8-14-2.

\uparrow perde os e^- da C. Val.

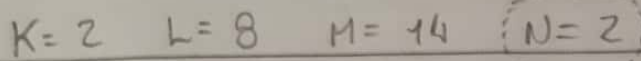
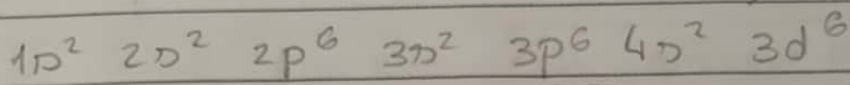
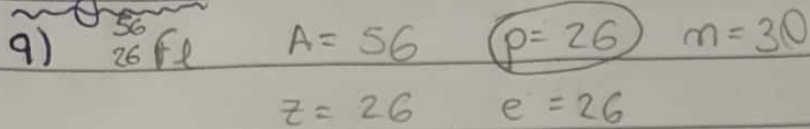
Depois perde aqui (átomo neutro)



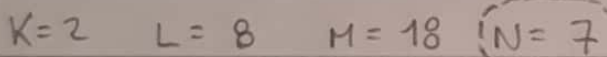
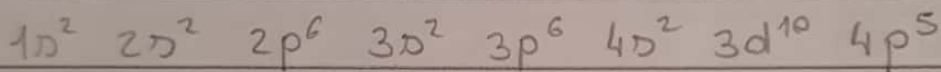
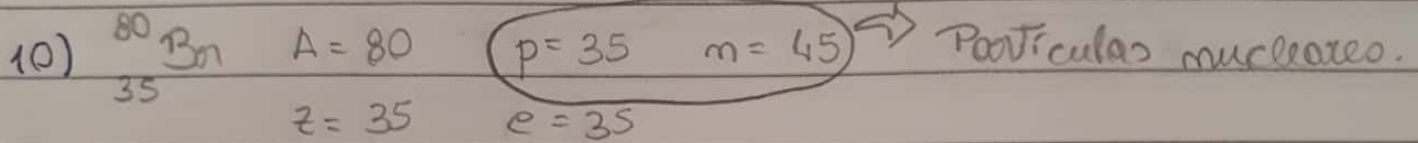
\rightarrow Ganhando $2e^-$ aqui na C. Val.



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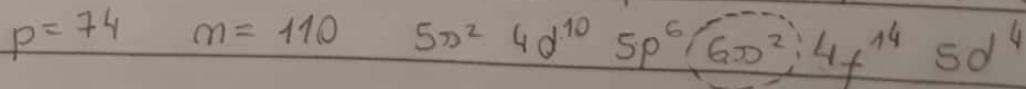
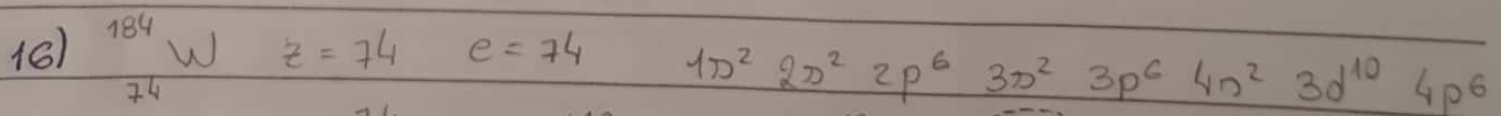
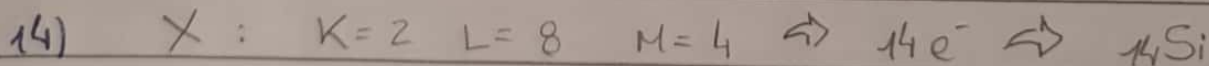
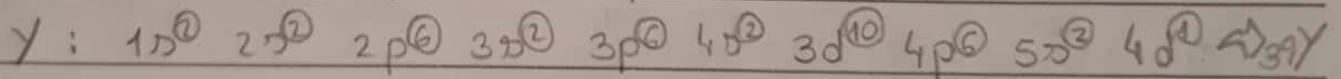
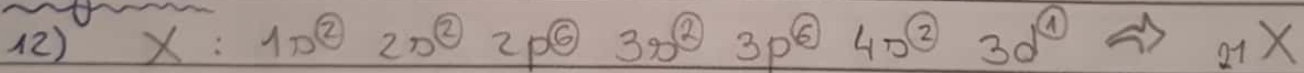


↪ Nível mais energético.

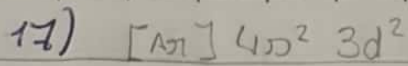


↪ Camada de valência: $7e^-$

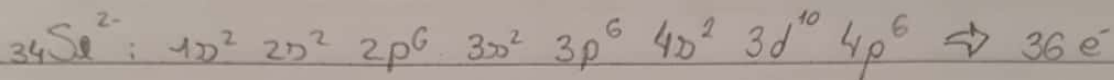
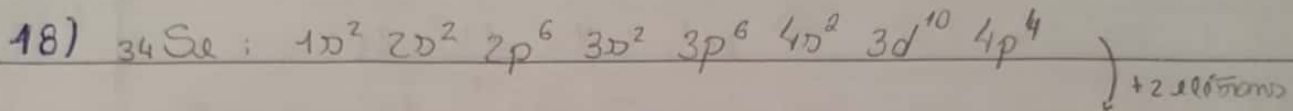
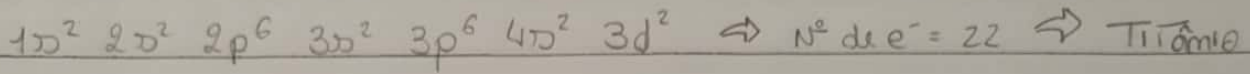
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↪ Camada de valência

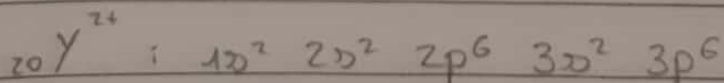
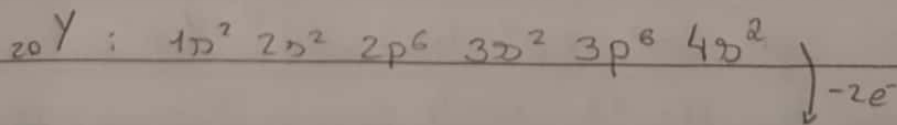
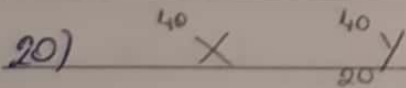
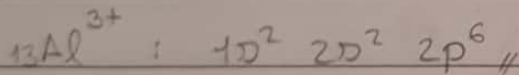
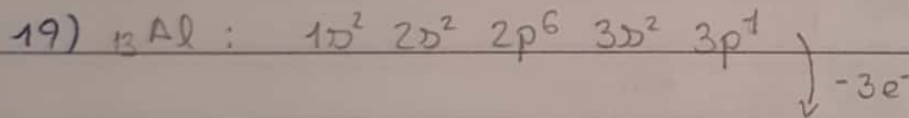


↪ Número atômica de gás Argônio

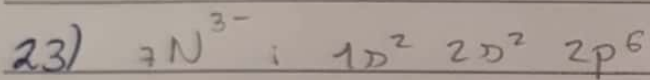
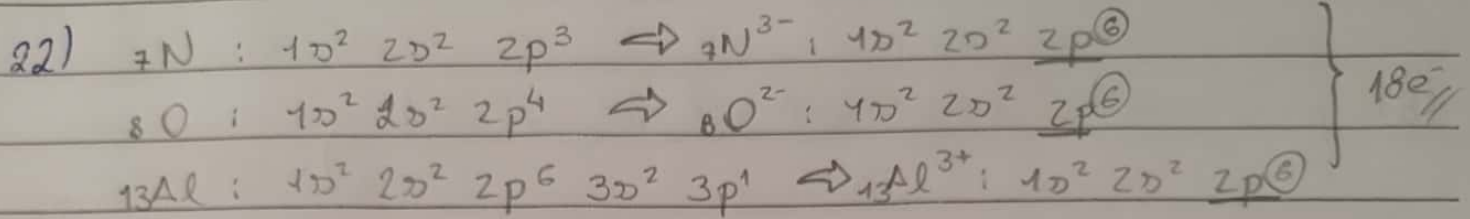
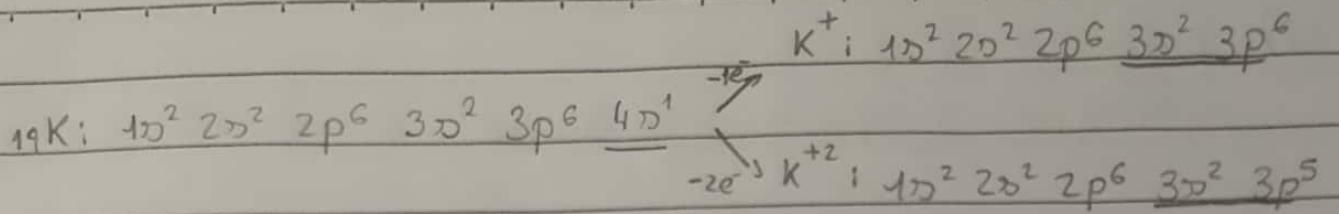


Então, 36 K π //

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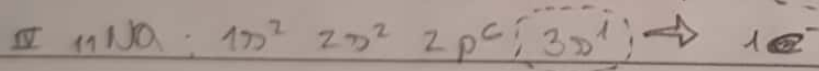
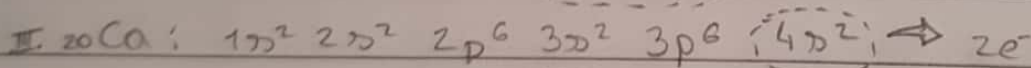
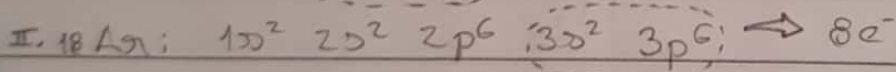
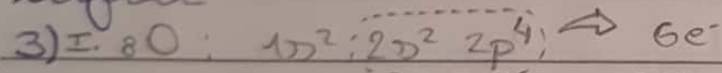


21)

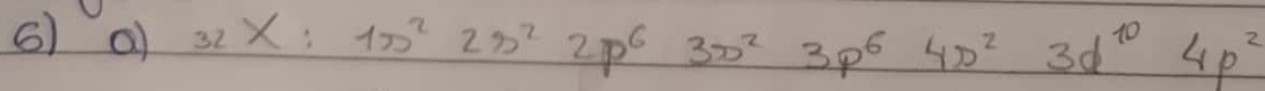


Apostila

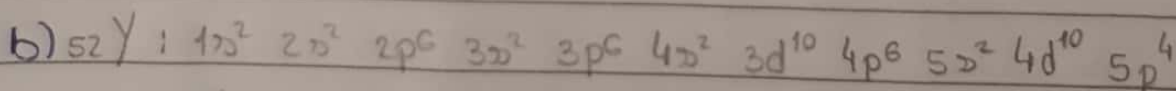
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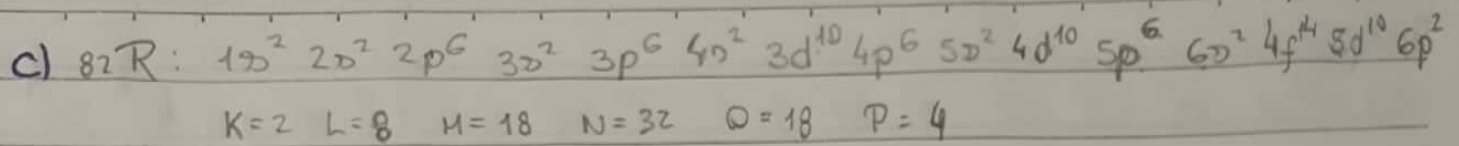
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$K=2 \quad L=8 \quad M=18 \quad N=4$

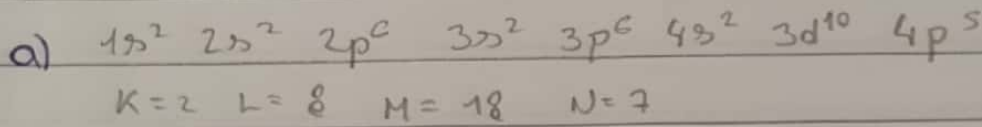


$K=2 \quad L=8 \quad M=18 \quad N=18 \quad O=6$



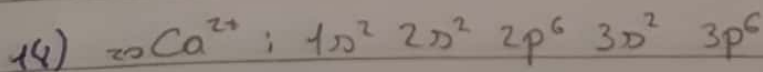
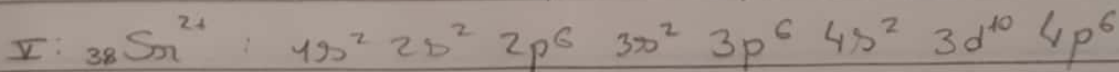
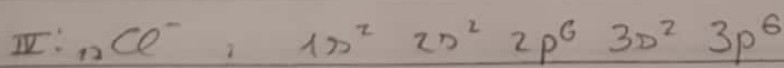
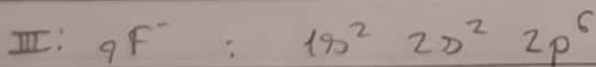
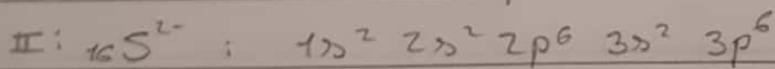
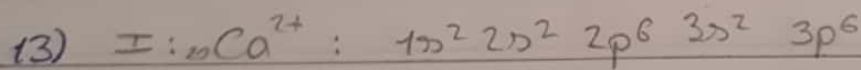
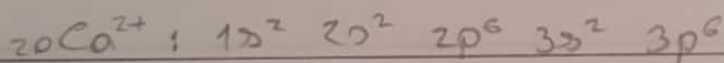
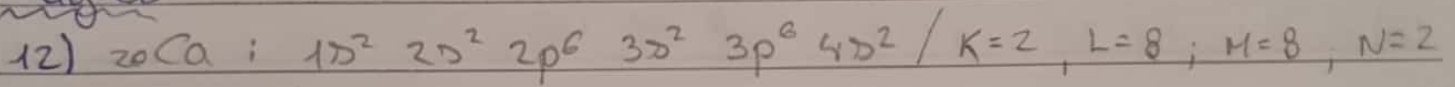
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7) $35X$



b) 7 elétrons

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